**Greenhouse Effect Lab**

Greenhouses allow gardeners to grow plants in cold weather. This is because the air inside the greenhouse stays warmer than the outside air. Short wavelength radiation from the sun passes through the glass, warming the interior of the greenhouse. The longer-wavelength radiation emitted does not pass through glass and is trapped in the greenhouse. This, along with the lack of mixing between the inside and outside air, keeps the greenhouse consistently warm.

Our atmosphere contains several chemical compounds that scientists refer to as greenhouse gases. The most common greenhouse gases include (in order of quantity): water vapor (H20), carbon dioxide (CO2), methane (CH4), nitrous oxide (N2O). These gases allow sunlight to enter the atmosphere freely. When sunlight strikes the Earth’s surface, some of it is reradiated back towards space as infrared radiation (heat) and some is absorbed by the Earth’s surface.

ES Exp 24 Fig1

Over time, the amount of energy sent from the Sun to the Earth’s surface should be about the same as the amount of energy radiated back into space, leaving the temperature of the Earth’s surface roughly constant.

Greenhouse gases absorb this infrared radiation and trap the heat in the atmosphere. The atmosphere acts as a natural blanket by preventing the Sun’s heat energy from radiating back into space, much like a greenhouse traps the Sun’s energy to warm someone’s plants even in the Earth’s surface by about 33oC. Without it, our planet would be too cold for humans to survive.

**MATERIALS:**

2 thermometers or temperature probes ruler

2 plastic bottles or Erlenmeyer flasks black paper (ground)

2 rubber corks with bore hole for thermometers

35 g baking soda + 65 mL vinegar (source of CO2)

65 mL water

lamp with 100 watt bulb (sun)

timer

Independent Variable: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Dependent Variable: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Constants: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Control: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Problem Statement: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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Hypothesis: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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**PROCEDURE**

1. Obtain two bottles and Label as Bottles 1 and 2.

2. Obtain 2 corks and gently insert thermometers into the corks. Thermometers must be 5 cm above the bottom of the container. During data collection the thermometers measure the “atmosphere” temperatures, not the liquid temperatures.

3. Add 65 mL of water to bottles 1.

4. Add 65 mL of vinegar to bottle 2.

5. Place each of the 2 bottles on the black paper directly in front of the lamp at distance of 25 cm.

6. Add 35 g baking soda to bottle #2, seal it with the cork/thermometer immediately, and swirl liquid for 10 seconds. Place bottle back into position on black paper in front of lamp.

7. Cover bottle #1 right afterwards with cork/thermometer.

8. Measure the starting temperature in each bottle.

9. Turn on the lamp.

10 Measure the temperature of each bottle every 2 minutes.

11. Turn off the lamp after 10 minutes.

12. Continue to measure the temperature of each bottle every 2 minutes.

13. At the 16 minute mark, measure the temperature of each bottle for the final time.

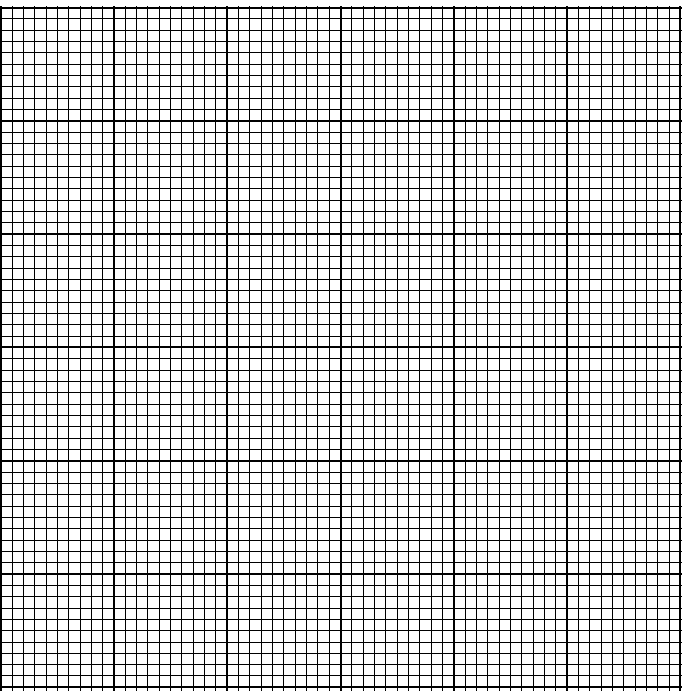
14. Calculate the Heat Increase as the difference between the start measurement and the 10 minute interval.

15. Calculate the Heat Retention as the difference between the 10 minute and 16 minute intervals.

16. Graph all the data.

**DATA**

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| Time | Bottle 1 (water) | | Bottle 2 (CO2) | |
| Trial 1 | Trial 2 | Trial 1 | Trial 2 |
| 0 |  |  |  |  |
| 2 |  |  |  |  |
| 4 |  |  |  |  |
| 6 |  |  |  |  |
| 8 |  |  |  |  |
| 10 |  |  |  |  |
| 12 |  |  |  |  |
| 14 |  |  |  |  |
| 16 |  |  |  |  |



What important greenhouse gases did each bottle contain?

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What is the Heat increase for each bottle when the light is on?

|  |  |  |  |
| --- | --- | --- | --- |
| Bottle | Trial 1 | Trial 2 | Average |
|  |  |  |  |
|  |  |  |  |

How much heat is retained by each bottle when the light is off?

|  |  |  |  |
| --- | --- | --- | --- |
| Bottle | Trial 1 | Trial 2 | Average |
|  |  |  |  |
|  |  |  |  |

Did either bottle heat up faster when the light was on? Explain.

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Did either bottle cool down faster when the light was off? Explain.

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Explain the scientific principals at work in this global warming lab:

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In what way(s) did the data **not** perfectly demonstrate the expected effect?

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